

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

GCE Advanced Subsidiary Level and GCE Advanced Level

**MARK SCHEME for the October/November 2011 question paper
for the guidance of teachers**

9701 CHEMISTRY

9701/33

Paper 3 (Advanced Practical Skills 1),
maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- Cambridge will not enter into discussions or correspondence in connection with these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2011 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – October/November 2011	9701	33

Question	Sections	Indicative material	Mark	
1 (a)	PDO Recording	I Thermometer readings for all experiments recorded to 0.0 or 0.5°C. (At least one recorded to 0.5°C.)	1	[7]
	ACE Interpretation	II Calculation of all temperature changes correct.	1	
	MMO Quality	Award III for a temperature rise followed by constant temperature (within 0.5°C).	1	
		Award IV and V for a maximum rise within 0.5°C of supervisor.	1	
		Award IV for a maximum rise within 1.0°C of supervisor.	1	
		Award VI and VII for the experiment 3 temperature rise within 0.5°C of supervisor. Award VI for the experiment 3 temperature rise within 1.0°C of supervisor.	1 1	
(b)	PDO Layout	I Axes correct and labelled: temperature change/ T change/ ΔT and volume/vol/V (of) sodium hydroxide/NaOH/ FA 1 and correct units /°C or (°C) or 'in °C'; /cm ³ or (cm ³) (allow NaOH in cm ³)	1	[4]
		II Scales chosen so that graph occupies at least half the available length for x- and y-axes.	1	
		III Plotting – all points accurate to within half a small square and in the correct square.	1	
		IV Draws two straight lines of best fit which intersect.	1	
(c)	ACE Interpretation	Reads to nearest ½ square to 1 or 2 dp volume of FA 1 and temperature rise from intercept. Do not award if ΔT at intercept (or point) < max ΔT from table unless candidate has clearly indicated the max ΔT is anomalous.	1	[1]

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – October/November 2011	9701	33

(d)	ACE Conclusions	I	The temperature/temperature change increases as more reaction/more hydrochloric acid/sodium hydroxide reacts/as more water formed.	1	[2]
		II	The temperature/temperature change stays constant/decreases when all acid/limiting reagent has reacted/excess NaOH is added.	1	
(e)	ACE Interpretation	I	Volume used in calculation is 65 cm ³	1	[2]
		II	Heat energy change calculated using candidate's value for ΔT correct to 3 or 4 sf	1	
(f)	ACE Interpretation		$\frac{25 \times 2}{1000} = 0.05$	1	[1]
(g)	ACE Interpretation	I	<u>Candidate's answer to (e)</u> Candidate's answer to (f)	1	[2]
	PDO Display	II	Correct calculation, conversion J to kJ and negative sign to 3 or 4 sf	1	
(h)	ACE Conclusions		So that rise in temperature is proportional to increase in energy produced/change in volume gives different change in temperature for same energy produced/increase in volume requires increase in energy for same temperature rise.	1	[1]
(i)	PDO Display	I	Number moles NaOH = number moles HCl (stated or clearly shown)	1	[2]
	ACE Interpretation	II	Calculates or expression for Concentration = $\frac{0.05 \text{ (ecf from (f))}}{\text{answer to (c)}/1000}$ If answer only, award mark if correct to 3 or 4 sf	1	
(j)	ACE Improvements		Use more concentrated solutions. (allow use ≤ 5 cm ³ water each time) Ignore all references to heat energy losses.	1	[1]
(k)	ACE Conclusions	I	Two straight intersecting lines (positive followed by zero gradient).	1	[2]
		II	Same ΔT and V shown as in (b).	1	
				[Total: 25]	

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
	GCE AS/A LEVEL – October/November 2011	9701	33

FA 3 is Na ₂ S ₂ O ₅ (s); FA 4 is Na ₂ CO ₃ (s); FA 5 is Na ₂ SO ₄ (s); FA 6 is Pb(NO ₃) ₂ (s) and (aq)					
2	(a)	MMO Decisions	(i) I Any named mineral acid or formula or (acidified) potassium dichromate Do not allow any reagent suitable for testing cations or more than one reagent.	1	[6]
		PDO Recording	(ii) II Tabulates evidence of 3 tests carried out with no repeat headings. Only consider observations with acid or dichromate.	1	
		MMO Collection	III Bubbles/effervescence in FA 4 .	1	
		MMO Decisions	IV Slower effervescence in FA 3 than FA 4 or FA 3 turns green and FA 5 stays orange if dichromate used.	1	
		ACE Conclusions	V Appropriate test with positive result used to test for either gas. VI All three ions correct from suitable observations. FA3 is a sulfite. FA4 is a carbonate. FA5 is a sulfate. (or correct formulae)	1	
	(b)	MMO Collection	(i) I FA 4 + FA 6 white ppt and FA 5 + FA 6 white ppt. II FA 6 + NaOH white ppt, soluble in excess sodium hydroxide. III Brown gas IV Gas relights glowing splint. V Yellow residue or crackling/decrepitating. VI Gas identified as oxygen or as NO ₂ from observations.	1	[6]
		ACE Conclusions	(ii) Lead/Pb ²⁺ provided correct observations with FA 6 + NaOH and FA 6 + FA 5 (sulfate).	1	[1]
		MMO Decisions	(iii) I Add HCl / H ₂ SO ₄ / KI / K ₂ CrO ₄ / NH ₃ *	1	[2]
		MMO Collection	II white ppt/white ppt/yellow ppt/yellow ppt/white ppt insoluble in excess. * If not Pb ²⁺ in (ii) but one of Al ³⁺ , Ba ²⁺ , Ca ²⁺ , Zn ²⁺ allow suitable reagent mark: K ₂ CrO ₄ for Ba ²⁺ and NH ₃ for the other three. However, observation must be correct for Pb²⁺ .	1	
					[Total: 15]